

1. (40 pts) 500 kg/hr of a feed solution containing a solute (C) mass fraction of 0.25 is to be extracted using 125 kg/hr of a recycled solvent stream (solute C mass fraction = 0.04, solvent mass fraction B = 0.95). The exiting raffinate shall be 0.10 mass fraction solute on a solvent free basis. Use the following equilibria data and phase diagrams on next pages to determine how many extraction stages are required. *Use the McCabe-Thiele method to determine the required number of stages.*

Diluent Rich (Raffinate)		Solvent Rich (Extract)	
x (x _B)	y (x _C)	x (x _B)	y (x _C)
0.07	0.22	0.30	0.42
0.06	0.17	0.40	0.39
0.045	0.12	0.52	0.30
0.04	0.06	0.64	0.18
0.038	0.04	0.685	0.12
0.035	0.02	0.73	0.06

2. (40 pts) Benzene will be stripped from a valuable oil by countercurrent contact with air in a tower packed with 2.0" ceramic Raschig rings. The contaminated oil (composition 98 mole % oil and 2 mole % benzene) will enter the tower at 2500 mol/hr and 95% of the entering benzene is to be removed. The flow rate of the incoming air will be 37,500 mol/hr. The density and viscosity of the dilute oil/benzene solution are well approximated by the properties of pure oil. The vapor phase behaves ideally. The tower will operate isothermally at 25 C and at a total pressure of 1 atm. The tower diameter shall be determined to give $\Delta P/\text{ft}$ of packing equal to 50% that of $\Delta P_{\text{flood}}/\text{ft}$. The equilibrium curve at these conditions is $y = 0.125 x$. You may regard the operating line as being straight. Calculation of the flooding velocity should be based on flow rates at the **top** of the tower, where they are largest. The density of the vapor at the top of the column can be approximated as that of air.

Data:

Air:

$$\text{MW} = 28.9$$

Benzene:

$$\text{MW} = 78.11$$

$$\text{Sc} = 1.76 \text{ in air}$$

$$\text{Sc} = 3500 \text{ in oil}$$

Oil:

$$\text{MW} = 106$$

$$\rho = 0.83 \text{ g/cm}^3$$

$$\nu = 2.84 \text{ cSt}$$

$$\mu = 2.36 \text{ cP}$$

Packing:

Table included in attachments

Gas Constant = $0.73024 \text{ ft}^3 \text{ atm} / (\text{R lbmol})$ R is degrees Rankine

Gas Constant = $82.05745 \text{ cm}^3 \text{ atm} / (\text{K mol})$

- What is the required diameter for the tower? Use the following graph and the correlation $\Delta P_{\text{flood}}/\text{ft} = 0.115 F_p^{0.7}$ inches water column per foot of packing
- What height of packing is required? Base your solution on $y - y^*$. Use the correlations from lecture to determine H_x and H_y .

$$H_{Oy} = H_y + m \frac{V}{L} H_x$$

3. (20 pts) *n*-Heptane undergoes mass transfer from a bulk gas (air + *n*-heptane) phase, where its mole fraction $y = 0.03$, to a bulk liquid (mineral oil + *n*-heptane) phase, where its mole fraction $x = 0.005$, through a gas-liquid interface. Temperature and pressure are 35 C and 1.0 atmosphere. The vapor pressure of *n*-Heptane at this temperature is 74.02 mm Hg. Mass transfer coefficients are as follows:

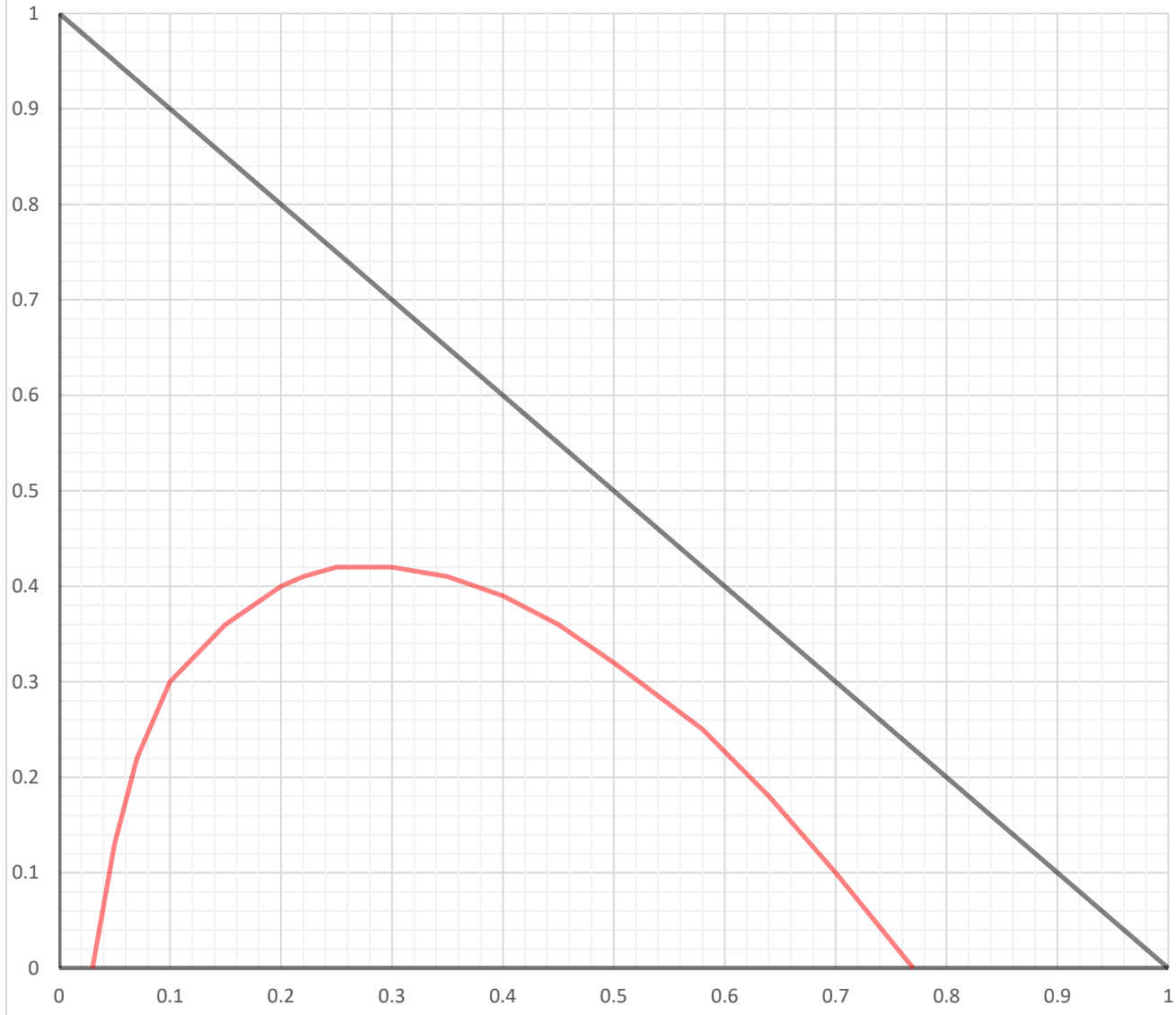
$$k_y = 7.0 * 10^{-6} \frac{mol}{cm^2 s}$$

$$k_x = 3.5 * 10^{-6} \frac{mol}{cm^2 s}$$

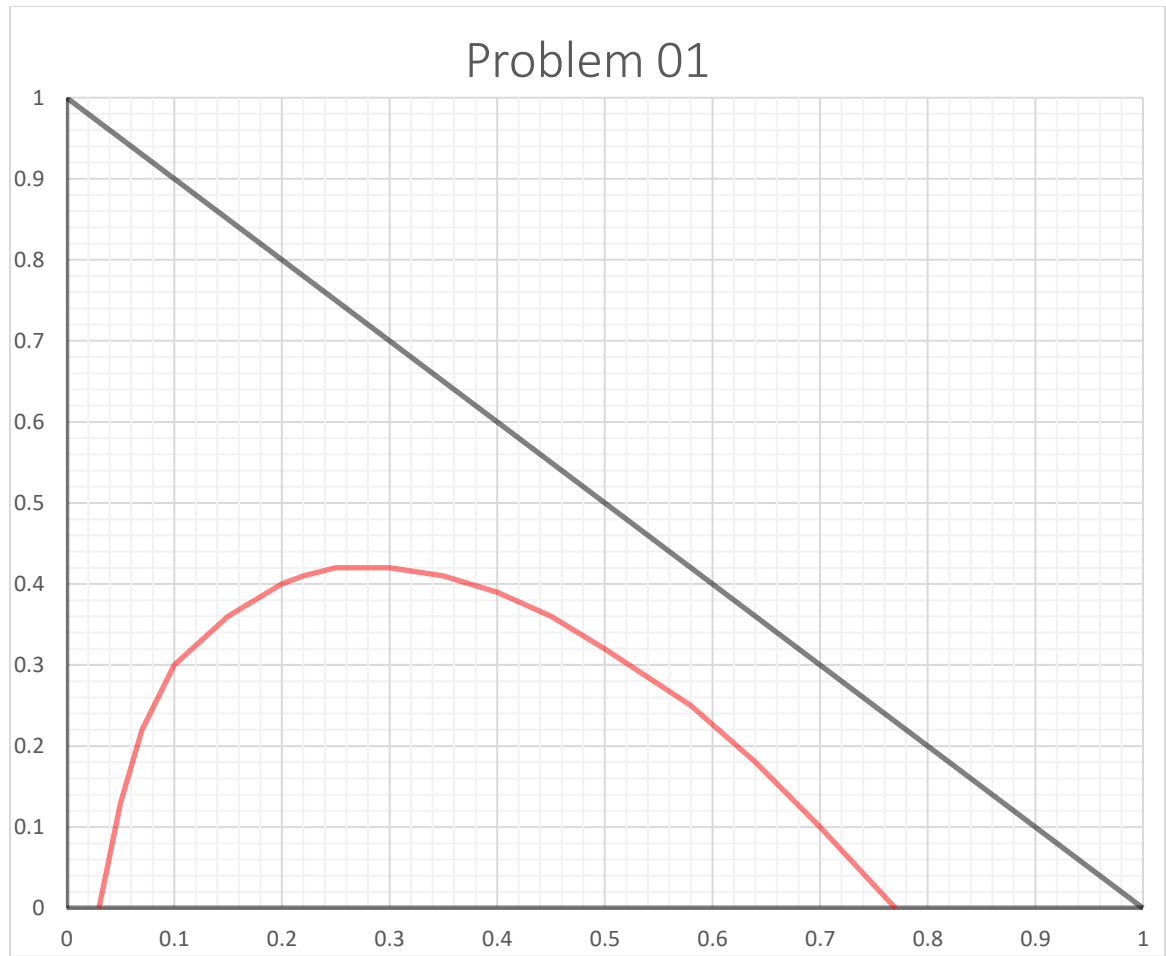
Assuming validity of Raoult's Law for the equilibrium relation,

- What are the overall mass transfer coefficients K_y and K_x ?
- What is the molar flux of *n*-heptane from gas to liquid?

Problem 01



Problem 01



Problem 1

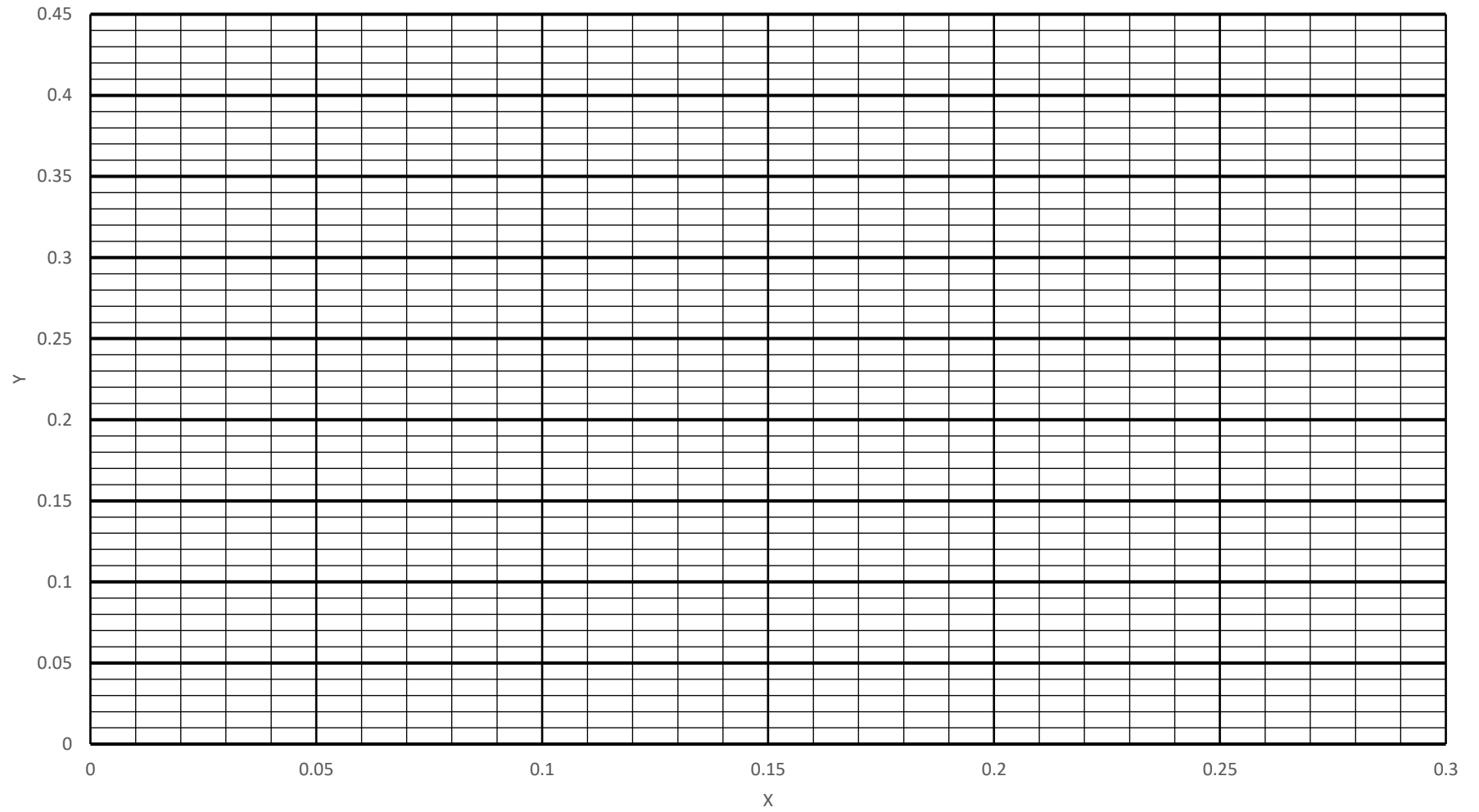


TABLE 18.1
Characteristics of dumped tower packings^{12,150,27}

Type	Material	Nominal size, in.	Bulk density, [†] lb/ft ³	Total area, [†] ft ² /ft ³	Porosity ϵ	Packing factors [†]	
						F_p	f_p
Raschig rings	Ceramic	1	55	112	0.64	580	1.52§
		1	42	58	0.74	155	1.36§
		1	43	37	0.73	95	1.0
		2	41	28	0.74	65	0.92§
Pall rings	Metal	1	30	63	0.94	56	1.54
		1	24	39	0.95	40	1.36
	Plastic	2	22	31	0.96	27	1.09
		1	5.5	63	0.90	55	1.36
Berl saddles	Ceramic	1	4.8	39	0.91	40	1.18
		1	54	142	0.62	240	1.58§
		1	45	76	0.68	110	1.36§
Intalox saddles	Ceramic	1	40	46	0.71	65	1.07§
		1	46	190	0.71	200	2.27
		1	42	78	0.73	92	1.54
		1	39	59	0.76	52	1.18
		2	38	36	0.76	40	1.0
		3	36	28	0.79	22	0.64

